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CHAPTER 9 REVIEW Stoichiometry SECTION 3

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g)

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mole is defined as the amount of a substance in 6.022×10^{23} (Avogadro's number) particles.

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Problems □ Do section review problems #1 through #4 on
page 301 of the textbook. Section Review page 301 1- The
branch of chemistry that deals with mass relationships in
compounds and in chemical reactions. $2 \text{ HgO} \rightarrow 2 \text{ Hg} + \text{O}_2$ 2-
a) 2 mol HgO 2 mol Hg & 2 mol HgO 1 mol O₂...

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SHORT ANSWER Answer the following questions in the space provided. 1. a.

6 Chemical Bonding

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